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Europäisches Patentamt
European Patent Office
Office européen des brevets

11 Publication number:

**0 355 218
A1**

12

EUROPEAN PATENT APPLICATION

21 Application number: **88307572.3**

51 Int. Cl.4: **C07C 1/04**

22 Date of filing: **16.08.88**

43 Date of publication of application:
28.02.90 Bulletin 90/09

84 Designated Contracting States:
DE FR GB NL

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54 **Process for catalytic hydrocarbon synthesis from CO and H₂ over metallic cobalt.**

57 The addition of water into a Fischer-Tropsch hydrocarbon synthesis reaction zone employing a catalyst comprising cobalt in a reduced, metallic form results in increased CO conversion and C₅ + hydrocarbon production, and a decrease in methane production.

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PROCESS FOR CATALYTIC HYDROCARBON SYNTHESIS FROM CO AND H₂ OVER METALLIC COBALTBACKGROUND OF THE INVENTIONField of the Invention

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This invention relates to a water addition process for increasing the activity of a Fischer-Tropsch, fixed bed, hydrocarbon synthesis reaction over a catalyst comprising metallic cobalt. More particularly, this invention relates to decreasing the methane make and increasing the CO conversion activity and C₅⁺ hydrocarbon selectivity of a Fischer-Tropsch hydrocarbon synthesis process which comprises co-feeding a mixture of H₂, CO and H₂O into a reaction zone containing a catalyst comprising cobalt in the reduced, metallic form wherein said mixture contacts said catalyst at elevated temperature for a time sufficient to convert at least a portion of said feed to C₅⁺ hydrocarbons.

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15 Background of the Disclosure

The production of hydrocarbons from mixtures of H₂ and CO via the Fischer-Tropsch process is well known to those skilled in the art. As opposed to the well-known "methanization" process which produces methane as synthetic natural gas from mixtures of H₂ and CO, the Fischer-Tropsch process is more generally aimed at producing higher value products such as chemical feedstocks and liquid fuels. Thus, high methane make is undesirable in Fischer-Tropsch synthesis processes because it is a relatively low value product which is formed at the expense of more desirable products. It is also uneconomical to try to convert the so-formed methane back into a CO and H₂ mixture and recycle it back into the reactor.

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Methane make in Fischer-Tropsch reactions is often expressed by a term known as methane selectivity. Methane selectivity can be defined by either of two methods. They are, (a) mole % CH₄ produced based on the amount of CO consumed or (b) weight % of CH₄ produced based on total hydrocarbon products formed.

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Many different catalysts and processes have been disclosed for Fischer-Tropsch synthesis, some of which have extremely high methane make. Thus, U.S. Patent 4,077,995 discloses synthesis of C₁-C₄ aliphatic hydrocarbons over a catalyst comprising a sulfided mixture of CoO, Al₂O₃ and ZnO while U.S. Patent 4,039,302 discloses C₁-C₃ hydrocarbon production using a mixture of the oxides of Co, Al, Zn and Mo. U.S. Patent 4,151,190 discloses producing C₂-C₄ hydrocarbons from mixtures of CO and H₂ using a supported catalyst comprising a metal oxide or sulfide of Mo, W, Re, Ru, Ni or Pt plus an alkali or alkaline earth metal, with Mo-K on carbon being preferred. U.S. Patent Nos. 4,243,553 and 4,243,554 disclose MoS₂ as a Fischer-Tropsch catalyst. Many other catalysts are known to be useful for Fischer-Tropsch synthesis employing metals such as iron, copper, titania, etc. These are known to those skilled in the art.

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The type of catalyst used and process conditions employed have an important bearing on CH₄ selectivity. For example, nickel gives a high CH₄ selectivity and is used mainly as a methanization catalyst. Methane selectivity usually increases with increasing temperature, decreasing pressure and increasing the H₂/CO ratio of the feed. Accordingly, process conditions are selected so as to minimize CH₄ selectivity while maintaining a relatively high reaction rate as is well known to those skilled in the art.

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It is known that CH₄ selectivity is influenced by the choice of promoter and support, such as alkali metal promoters reducing CH₄ selectivities of iron catalysts. It is also known in the art that noble metals such as ruthenium supported on inorganic refractory oxide supports exhibit superior hydrocarbon synthesis characteristics with relatively low methane production. Thus, U.S. Patent No. 4,088,671 suggests minimizing methane production by using a small amount of Ru on a cobalt catalyst. Examples of supported ruthenium catalysts suitable for hydrocarbon synthesis via Fischer-Tropsch reactions are disclosed in U.S. Patent Nos. 4,042,614 and 4,171,320. It is also known that the type of support used also influences methane production. In the case of supported ruthenium catalysts, the use of a titania or titania containing support will result in lower methane production than, for example a silica, alumina or manganese oxide support.

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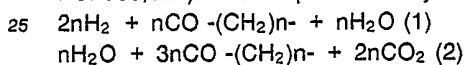
European patent application 83201557.2 relates to a Fischer-Tropsch hydrocarbon synthesis process employing, as a catalyst, cobalt supported on silica promoted with zirconium, titanium or chromium wherein the cobalt is deposited on the silica carrier by a kneading process. In this process, 10 to 40 volume percent steam is added to the feed, based on the H₂/CO/H₂O mixture. The amount of cobalt present in the catalyst is 10-40 pbw based on 100 pbw of silica. In the Examples, very little enhancement in either CO conversion

or C₃⁺ selectivity is shown as a result of the water addition to the feed. U.K. patent application 2146350A relates to a Fischer-Tropsch hydrocarbon synthesis process employing a similar catalyst, followed by a hydrocracking process wherein the Fischer-Tropsch products are hydrocracked over a noble metal catalyst. The same type of catalyst is also used in the Fischer-Tropsch processes described in U.K. patent application nos. 2,149,812A and 2,149,813A.

U.S. Patent 3,927,999 relates to a process for producing a methane rich fuel gas containing 70 to 98 mole % methane by contacting a mixture of H₂, CO and H₂O with a suitable catalyst at elevated temperature. This invention is based on the discovery that the methane content of the product gas from the methanator is maximized by adjusting the mole % H₂O in the synthesis gas feed to a critical value in the range of 1 to 3, while maintaining the H₂/CO mole ratio of the synthesis gas feed to a critical value in the range of about 1 to 1.15. The feed gas may also contain methane. Group VIII transition metals such as iron, nickel and cobalt are suggested as suitable catalysts, but only nickel on alumina is actually used. Example 1 shows the production of a fuel gas containing 96.4 mole % CH₄ from a feed gas containing 48.1 mole % CH₄ over a catalyst containing nickel, thoria, magnesia and kieselguhr.

U.S. Patent 2,479,439 relates to a process for alternately increasing the activity of an alkali metal promoted, powdered iron catalyst or decreasing the coke buildup thereon, or accomplishing both simultaneously by passing an aqueous solution of alkali metal salt, such as KF, into the hydrocarbon synthesis reaction zone to contact the iron catalyst. The presence of the water is said to remove carbon from the iron catalyst. U.S. Patent 2,539,847 relates to a Fischer-Tropsch hydrocarbon synthesis process employing a catalyst consisting of thoria promoted cobalt supported on bentonite. U.S. Patent 4,568,663 relates to a Fischer-Tropsch hydrocarbon synthesis process employing a rhenium cobalt on titania catalyst.

The Fischer-Tropsch (FT) process, the process to which this invention relates, must also be distinguished from the well known Kolbel-Engelhardt (KE) process (described, for example, in Canadian Patent No. 530,932). The FT process may be described by equation (1), the KE process by equation (2).



The two processes may appear similar in that they yield similar products and can employ Group VIII metals as catalysts. The reactions are not similar because their reaction paths differ in the formation of hydrocarbons and because FT yields H₂O as a product while KE consumes H₂O as a reactant. The reactions can be distinguished in that in FT there is no net consumption of H₂O, whereas in KE there is net consumption of H₂O. Even if hydrogen is added to the reaction mixture for KE, water will be consumed.

There exists a need in the art for Fischer-Tropsch processes useful for the conversion of mixtures of CO and hydrogen to C₅ + hydrocarbons at high CO conversion levels, and high hydrocarbon yields, with relatively low methane make.

SUMMARY OF THE INVENTION

It has now been discovered that in a fixed bed, Fischer-Tropsch hydrocarbon synthesis reaction for producing C₅ + hydrocarbons from a gaseous feed mixture of H₂ and CO in the presence of a catalyst comprising cobalt, one can increase both the CO conversion activity and the C₅ + hydrocarbon selectivity and, at the same time, decrease the methane make by adding H₂O to the reaction zone. Thus, the present invention relates to a Fischer-Tropsch process for synthesizing C₅ + hydrocarbons by introducing into a catalytic reaction zone a feed mixture of CO, H₂, and H₂O wherein said feed contacts a catalyst comprising cobalt at elevated temperature and for a time sufficient to convert at least a portion of said feed to C₅ + hydrocarbons. By cofeeding H₂O into the reaction zone, along with the H₂ and CO, it has been found that C₅ + hydrocarbon production is increased, CO conversion is increased and CH₄ production is decreased. By cobalt is meant cobalt in the reduced, metallic form, preferably a high surface area cobalt like cobalt black.

The H₂O that is added to the reaction zone may be in the form of steam or moisture or a suitable H₂O precursor, such as C₁-C₆ alcohols, for forming H₂O in-situ in the reaction zone. It is essential to the understanding of the process of this invention that the H₂O introduced into the reaction zone is external H₂O and not that H₂O which is formed in-situ in the reaction zone as a consequence of the Fischer-Tropsch hydrocarbon synthesis reaction from the H₂ and CO. It has also been found that the process of this invention improves with increasing pressure in the reaction zone and with decreasing CO conversion.

In a preferred embodiment of the invention, the synthesis gas feed, carbon monoxide and hydrogen is substantially free of and preferably completely (except for trace amounts) free of methane or other light hydrocarbons. The hydrocarbon free synthesis gas feed is easily accomplished in a once-through system

where there is no recycle of unconverted hydrogen and carbon monoxide (along with undesirable light hydrocarbons) directly to the hydrocarbon synthesis reaction zone. Preferably, the synthesis gas feed contains only H₂, CO and water and is substantially free of light hydrocarbons, e.g., methane. Allowable limits on light hydrocarbons in the synthesis gas feed are no more than about 1 vol.% and preferably less than about 0.5 vol.%.

DETAILED DESCRIPTION

The process of the present invention resides in adding, to the Fischer-Tropsch reaction zone, H₂O or a suitable H₂O precursor such as an alcohol. The amount of H₂O added to the reaction zone will range from about 1-70 volume % of the total feed mixture of H₂O, CO and H₂ and, preferably, from about 5-30 volume %. As the extent of CO conversion increases, the amount of water produced in-situ in the reaction zone increases and, concomitantly the beneficial effect of introducing additional water into the reaction zone to increase CO conversion, reduce CH₄ selectivity and increase C₅ + hydrocarbons selectivity decreases. Thus, in some cases, it may be advantageous to practice the process of this invention CO conversion levels below about 60 percent. That is, below about 60% per pass, reaction zone or stage. As the pressure in the reaction zone increases, the beneficial effect of the process of this invention of adding water to the reaction zone increases with respect to increasing CO conversion activity, decreased CH₄ selectivity and increased C₅ + hydrocarbon selectivity. At relatively low pressures in the reaction zone (i.e.; less than about 1 atmosphere), little effect will be seen in increased conversion activity, etc. by adding H₂O to the reaction zone. Thus, the process of this invention will be operated at a pressure above about one atmosphere. In general the pressure will range from about 1-50 atmospheres, and more preferably 5-30 atmospheres.

In the process of this invention, the addition of H₂O enhances the formation of C₅ + hydrocarbons and reduces CH₄ make in a uni-directional or linear fashion. It has also been found that the rate of the Fischer-Tropsch reaction employing the process of this invention increases with increasing partial pressure of H₂O at any fixed total pressure in the reaction zone, up to a point, after which point the rate slowly goes down with continually increasing H₂O partial pressure.

In general, the Fischer-Tropsch hydrocarbon synthesis reaction process of this invention is carried out at a H₂:CO mole ratio of greater than about 0.5, and preferably the H₂:CO mole ratio ranges from about 0.5 to about 6, more preferably from about 0.5 to about 3, at gas hourly space velocities ranging from about 100 V/Hr/V to about 5000 V/Hr/V, preferably from about 300 V/Hr/V to about 1500 V/Hr/V, at temperatures ranging from about 150° C to about 300° C, preferably from about 180° C to about 240° C, and pressures above about 1 atm., preferably ranging from about 1 atm. to about 50 atm., more preferably from about 5 atm. to about 40 atm. and still more preferably from about 5-30 atmospheres.

As previously stated, the hydrocarbon synthesis process of this invention employs a catalyst comprising cobalt in its reduced, metallic-form, preferably a high surface area cobalt like cobalt black. The cobalt may be in the form of a finely divided powder, or granules which may be mixed with a suitable diluent to aid in heat transfer and removal from the reaction zone. The cobalt may be supported on a suitable support material, such as cobalt plated on carrier metal. However, this is not meant to include cobalt dispersed on inorganic refractory oxide supports. The cobalt in bulk form, or plated, explosion-coated, etc. may also be in the form of various high surface area shapes such as spirals, metal wool, honeycomb configurations, etc., the choice being left to the practitioner.

When supported catalysts are employed, for example, cobalt on a support such as titania, silica, alumina, or silica-alumina, the improvement in the process from the addition of water is apparent only with supports of a relatively low surface area. Thus, supported cobalt catalyst preferably have a surface area of less than about 40 square meters per gram of catalyst (BET).

The invention will be more readily understood by reference to the examples below.

EXAMPLES

Example 1 12% Co/TiO₂

A bulk cobalt catalyst, designated as a cobalt black catalyst, was prepared by a conventional method known in the art, i.e., adding a stoichiometric amount of ammonium carbonate (as an aqueous solution) to an aqueous solution of cobalt nitrate, filtering the precipitate, washing the filtered solid with deionized water,

drying at 120 °C, and calcining at 500 °C in air for 5 hours. The resulting material showed an X-ray diffraction pattern of Co_3O_4 which, when reduced at 450 °C in a H_2 stream, produced a cobalt black catalyst having a surface area of 7.7 m^2/gr .

5 The performance test of this catalyst was carried out by charging an intimate mixture of the Co_3O_4 powder (6.8 grams) and a diluent (quartz powder, 80-140 mesh, 18 grams) into a down-flow fixed bed reactor made of 3/8" OD stainless steel tube with a concentric 1/8" OD thermocouple well. When reduced, 6.8 grams of Co_3O_4 will yield 5.0 grams of cobalt black. The use of diluent and an aluminum block jacket fitted tightly around the reactor minimized uneven temperature profile along the bed. The Co_3O_4 was then re-reduced in-situ in the reactor in a flowing H_2 stream (200 cc/m, 1 atm) overnight at 450 °C, cooled to 10 175 °C and then pressurized to 20 atm using a pre-mixed gas composed of 63.9% H_2 , 32.1% CO and 4.0% N_2 . At pressure and at a preset flow rate, the temperature was raised to 200 °C over a period of one hour and then data acquisition was initiated. The rate of CO conversion and the rates of various hydrocarbon products formation were monitored using two on-line GC's. The data taken after 70 hour on-stream-time are shown in Table 1.

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TABLE 1

EFFECTS OF EXTERNAL H₂O
ADDITION ON ACTIVITY
SELECTIVITIES OF CO BLACK

Temperature = 200°C

Feed Gas Composition = 63.1% H₂/33.0% CO/3.9% N₂

SV = 3600 Scm³/g hr (calcined for H₂ and CO)*1

	<u>Run 1</u>	<u>Run 2</u>	<u>Run 3</u>	<u>Run</u>
<u>4</u>				
% H ₂ O Added*2	0	12.5	27.7	50.0
Total Pressure, atm	20.7	21.6	24.9	28.6
PH ₂ + Pco, atm	20.0	18.7	18.9	18.6
PH ₂ O, atm	0	2.3	5.2	9.3
% CO Conv.	12.1	27.3	28.7	18.8
CH ₄ Select*3	10.5	7.1	5.8	4.0
CO ₂ Select*3	0.27	0.21	0.34	
Hydrocarbon Product Distribution, wt. %				
CH ₄	11.5	8.1	6.7	4.6
C ₂ -C ₄	15.6	8.7	7.8	8.6
C ₅ -C ₉	17.4	9.5	9.0	9.7
C ₁₀ ⁺	55.5	73.7	76.5	77.1
C ₁₀ -C ₂₀	21.7	17.0		
C ₂₁ -C ₃₀	21.7	13.8		
C ₃₁ -C ₄₀	6.9	11.1		
C ₄₁ -C ₅₀	4.7	8.1		
C ₅₁ ⁺	11.8	23.7		

*1 cc of H₂ and CO measured at 1 atm, 22°C per gram of catalyst per hour.

*2 Moles of H₂O added per 100 moles of CO and H₂

*3 Moles of CH₄ or CO₂ per 100 moles of CO converted

Comparison of the data with and without the external H₂O addition clearly establishes the advantages of H₂O addition; (1) the CO conversion increases markedly (up to 2.5 fold increase), (2) the methane make is greatly reduced and, (3) the desired heavy hydrocarbon product yields, which may be gauged by the C₁₀ + selectivity, increase dramatically.

Thus, this example clearly demonstrates the benefits of conducting a Fischer-Tropsch hydrocarbon synthesis reaction with external H₂O addition in the presence of a catalyst comprising metal cobalt.

Example 2

Another batch of a bulk cobalt catalyst was prepared using highly purified Co(NO₃)₂•6H₂O (99.99%

pure Puratronic grade) according to a procedure similar to that described in Example 1. The evaluation procedure was also identical to that used in Example 1. After an extended reduction in a flowing H₂/He mixture at 500° C, the material was passivated and found to have about 1.5 m² surface area per gram. This catalyst (5 grams) and 10 grams of quartz powder were mixed and charged into the reactor and pretreated as described in Example 1. At 201.5° C and 20.7 atm., a gas mixture composed of 63.1% H₂, 33.0% CO and 3.9% N₂ was passed at a rate of 105 Scm³/m. After one day on stream the data in Table 2 were obtained.

TABLE 2

EFFECTS OF ADDED H₂O ON THE ACTIVITY AND SELECTIVITIES OF CO HYDROGENATION OVER A HIGH-PURITY BULK CO CATALYST

Temperature = 202° C
 Feed Gas Composition = 63.1% H₂/33% CO/3.9% N₂
 Pressure = 20 atm.
 SV = 1200 Scm³/g hr

	<u>Run 5</u>	<u>Run 6</u>
% H ₂ O Added	0	21
% CO Conversion	7.2	15.6
CH ₄ Selectivity	8.2	4.2
1-olefin/paraffin Ratio		
C ₂	0.13	0.36
C ₃	2.47	3.68
C ₄	1.87	2.52
Hydrocarbon Product Distribution, wt. %		
CH ₄	9.4	4.8
C ₂ -C ₄	21.1	20.4
C ₅ -C ₉	11.0	11.6
C ₁₀ +	58.5	63.2

Claims

1. A fixed bed catalytic process for synthesizing C₅ + hydrocarbons which comprises reacting, in a reaction zone at elevated temperatures, hydrogen and carbon monoxide substantially free of light hydrocarbons in the presence of added water, wherein the catalyst is comprised of cobalt, and wherein there is no net consumption of water.
2. The process of claim 1 wherein the water (or a precursor thereof) is added to the feed prior to entering the reaction zone.
3. The process of claim 1 wherein the water (or a precursor thereof) is added to the reaction zone.
4. The process of any one of claims 1 to 3 wherein the amount of water added is in the range of from 1 to 70 vol.% of the total amount of H₂O, CO and H₂.
5. The process of any one of claims 1 to 4 wherein the catalyst is comprised of cobalt on a support and the support has a surface area of less than about 40 m/gm.
6. The process of any one of claims 1 to 5 wherein the pressure in the reaction zone is greater than 1 atmosphere.



DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int. Cl. 4)
D,X	EP-A-0 109 702 (SHELL INTERNATIONAL RESEARCH MAATSCHAPPIJ N.V.) * Claims * ---	1-2,4-6	C 07 C 1/04
X	DE-C- 487 379 (I.G. FARBENINDUSTRIE) * Whole article * ---	1-2,4	
A	DE-C- 716 853 (STUDIEN- UND VERWERTUNGS-GESELLSCHAFT mbH) ---		
A	US-A-2 497 964 (S.D. SUMERFORD) ---		
A	NL-C- 77 441 (STANDARD OIL DEVELOPMENT CO.) -----		
			TECHNICAL FIELDS SEARCHED (Int. Cl.4)
			C 07 C 1/00
The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 16-03-1989	Examiner VAN GEYT J.J.A.
CATEGORY OF CITED DOCUMENTS		T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document	
X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document			