EXPERIMENTS ON THE REMOVAL OF BUTADIENE FROM THE BUTANE DEHYDROGENATION-GASES-BY-MEANS-OF-SELECTIVE-HYDROGENATION-

In order to reduce the acid consumption in sulphuric acid alkylation which is caused by the presence of diolefins, mainly of butadiene, in the butylene feed, experiments were made to remove the diolefins by selective hydrogenation.

While laboratory experiments were successful, both at atmospheric pressure and at a pressure of 147 psi., large-scale experiments which were only carried out at atmospheric pressure gave much less satisfactory results.

Laboratory Experiments.

The exit gas from the butane dehydrogenation reactor was processed in a laboratory apparatus consisting of a quartz tube with a capacity of 15.2 - 18.3 cubic inches. The first experiments were made with catalyst 4788 which is a chromium-nickel catalyst, used for the selective hydrogenation of C2H2 in ethylene in the plant.

The dehydrogenation gas contains 15 to 20% C₄H₈, 0.5 to 1.5% C₄H₆ and 15 to 20% H₂ and was hydrogenated without the addition of fresh hydrogen. The experiments were made in the temperature range from 302 - 572°F, at space velocities from 1,000 to 5,000 volumes of gas per volume of catalyst per hour. The results are shown in the attached tables. Table 1 gives the results obtained with gas washed with 10% NaOH solution, whereas the results given in Table 2 are obtained without caustic washing. The data indicate that the diplefins are hydrogenated to olefins and that only traces of the olefins have reacted. Tests in a special alkylation apparatus indicated that the acid consumption decreased from 25 to 30% for the untreated dehydrogenation gas to 15 to 20% after selective hydrogenation.

Table 3 shows the effect on acid consumption by hydrogenation under various conditions using catalyst 4788.

Test results with the commercial hydrogenation catalyst 3076 (WS2.2NiS) which is not sensitive to sulphur compounds are shown in Table 4. The particular catalyst used had been previously in plant service for several months (hydrogenation of heavy gas oil from Fischer-Tropsch synthesis).

Good results were also obtained with catalyst 7846 (WS2.2NiS on an alumina carrier) and catalyst 7058 (MoS2.2NiS) among others.

Raising the hydrogen concentration of the dehydrogenation gas to 50% by the addition of fresh hydrogen also gave favorable results.

The results of hydrogenation tests at a pressure slightly above atmospheric are given in Table 5 for several catalysts.

All laboratory results indicate that butadiene in the butane dehydrogenation gas can be selectively hydrogenated at atmospheric and slightly elevated pressure. The olefins are practically not changed under these conditions.

Plant Experiments.

On the basis of the laboratory results, experiments were carried out on a technical scale using catalysts 3076 and 4788. The experiments were made at atmospheric pressure, temperatures of 356 - 392°F, and space velocities of 1,000 to 4,000 volumes per volume of catalyst per hour. No satisfactory results were obtained under these conditions and only when the temperature was raised to 482 - 572°F, hydrogenation of butadiene was obtained, provided the space velocity was sufficiently low.

The dehydrogenation gas from the reactor was freed from dust and passed into a steam-heated preheater from where it flowed to two reactors connected in parallel. The dimensions of the reactors were as follows:

	Reactor 1	Reactor 2
Diameter	2.3 feet	2.6 feet
Height	7.2	7.4 "
Catalyst Bed	4.5 "	3. 9 "
Volume of Catalyst	17.7 ou. ft.	17.7 cu. ft.

The reactor was not heated but insulated against radiation losses. In cases where higher temperatures were desired, the steam-heated preheater was replaced by a gas-heated preheater.

All experiments were practically without success, with the exception of the experiment in which freshly-reduced catalyst 3076 was used at 570°F. and a space velocity of 500 volumes of gas per volume of catalyst per hour. The expected decrease in acid consumption in the alkylation was not reached because only part of the dehydrogenation gases could be passed through the experimental unit. The activity of the catalyst in the successful experiment decreased, however, considerably after 20 days. The results of the tests in the large-scale unit are given in Table 6.

It might be possible to prolong the activity of a catalyst by a preliminary purification of the dehydrogenation gas by means of coke filters. Because of air attacks no large-scale experiments under pressure could be made.

Table 1

	Reactor Temp.			Analyses			
Pressure	or.			Olefins		Diolefins	
psig.	top	center	bottom	reed	produce	1.000	-produce-
, o	266	446	482	9.8	9.4	0.7	0.2
0	239	401	500	12.3	12.0	0.5	0.2
.0	230-	392	493	14.4-	14.2	0.7_	0.4
c ·	212	383	473	16.8	17.1	0.8	0.1
0	266	473	572	. 17.2	17.5	1.1	0.1
	psig. 0 0 0 0	Pressure psig. top 0 266 0 239 0 230- 0 212	Pressure OF. psig. top center 0 266 446 0 239 401 0 230 392 C 212 383	psig. top center bottom O 266 446 482 O 239 401 500 O 230 392 491 C 212 383 473	Pressure OF. OI psig. top center bottom feed O 266 446 482 9.8 O 239 401 500 12.3 O 230 392 491 14.4 C 212 383 473 16.8	Pressure OF. Olefins psig. top center bottom feed product 0 266 446 482 9.8 9.4 0 239 401 500 12.3 12.0 0 230 392 491 14.4 14.2 C 212 383 473 16.8 17.1	Pressure OF. Olefins Dio psig. top center bottom feed product feed 0 266 446 482 9.8 9.4 0.7 0 239 401 500 12.3 12.0 0.5 0 230 392 491 14.4 14.2 0.7 0 212 383 473 16.8 17.1 0.8

Table 2

	•	R	eactor T	emp.		Analy			
Space Velocity Vol/Vol/hour	Pressure psig.	top	oF. center	bottom		efins product	Dio feed	lefins product	
2000	Ö	347	572	662	12.6	12.2	0.4	0.1	
3000	۵.`	248	446	581	13.4	13.5	0.9	0.3	
4600	0	266	527 .	662	14.4	15.0	0.4	0.1	
4500	0	320	536	662	11 .2	11.0	0.3	0.0	

Table 3

							nsumption
Ce	talyst	Temp.	Space Velocity Vol/Vol/hour	Pressure psig.	Caustic Wash	lbs.acid/ before hydrog.	obl.alkylate after hydrog.
	4788	-: 482	2000	0	no	73.6	58.5
	4788	482	2000	0	yes	76 .2	42.7
	4788	572	-3000	, 0	yes	76.2	39.8
	4788	572	3000	0	ya s	76 .2	35.6
•	4788	572	3000	0	no	76 .2	39.3
	4788	572	3000	0	no	76,2	-~ 30.7
	4788	572	4000	0	no	90.9	54.5

				Table	4			-	
Catalyst	Space Velocity Vol/Vol/hour	Temp.	Pressure palg.	% Olefins	ofins produot	% Dic	% Diolefins ed product	Acid Consum	Consumption
-	. ,				,			7 9 9	n n n n n n n n n n n n n n n n n n n
3076	2000	302	0	17.6	17.6	1.0	8°0	 06	37.3
#	4000	392	o ~ .	17.8	17.8	1,2	0.7	92°4	51.5
	4000	482	O	16.6	16.0	1.0	000	92.4	51.5
2	4000	572	•	17 .6	17.3	0.7	2°0	92°4	***************************************
	0009	573	0	20°5	20°0	1.1	0.7	106.8	39.8
5076 Freshly	4000	392	. 0	14.0	14.0	0.4	0°0	3, 721	31.2
reduced	4000	392	o	14.6	14,2	0.7	0°0	130.0	88 88 9
	4000	, 392	0	14.2	13.8	9°0	0.1	122.6	50 3
	4000	572	0	14 °6	14.2	9.0	0.5	122.6	30.0
		:		:		•			
					1				
	• .	,	١	Table	او				J
Catalyst	Space Velocity	Temp	Pressure	% 01	% Olefins	% D1	% Diolefins	Acid Co	Acid Consumption
	.yo1/vo1/tonr	op.	patgo	feed	product	feed	product	1bs.soid/	lbs.aoid/bbl-alkylate
	, .				•	· • · ·		feed	product
3076	3000	392	117.5	15.0	14.8		0°0	92°4	34.4
nseq	4000	485	117.5	14,2	14.0	8.0 8.0	4.0	92 °4	32°4
in plant	0009	482	117.5	16.6	15.4	0.7	4°0	92 ,4	8 8 8
#	0009	572	117.5	15.2	14.8	9°0	0.3	106.8	38,1
	8000	572	117.6	13.8	13.8	9°0	9°0	106 .8	32°6
3076 Freshly	4000	572	117.6	16.2	15.0	6.0	0.5	129.0	49.1
reduced 7846	2000	572	117.5	14,4	13.4	0.5	0.25	110,5	45 ,6
7058	2000	. 572	117.5	13.5	13.2	0.75	0.50	135,0	42.5

Reactor Filling	both reactors filled with 17.6 cu. ft. of 4788 each	reactor 1, 17.6 cu. ft. 3076 from Kogasın reactor, reactor 2, 21.2 cu. ft. of 3076 from plant.	reactor 1, 21.2 cu. ft. 4788 reactor 2,33.6 cu. ft. 3076	reactor 2, 26.5 cu. ft. 3076 freshly reduced.	reactor 1, 21.2 cu. ft. 4788 reactor 2, 33.6 cu. ft. 3076	reactor 1, 21.2 cu. ft. 4788 reactor 2, 33.6 cu. ft. 3076	reactor 1, 24.7 cu. ft. 3076 (0.16") Freshly reduced reactor 2, 33.6 cu. ft. old 3076.
Acid Consumption .acid/bbl.alkylate feed product	86.0 unchanged	=	2	E.	E		" " 110 ° 5
The	6 73.6-86.0	9 73.6	7 103	122.8	.5 44.2	44 44 44 44 44 44 5 6 6 6 6 6	110
uot 7 C4H6	0.3-0.6	0.5-0.9	0.4-0.7	0.25	0,25-0,5 0,25 0,25 0,25	0.85 0.35 0.35 0.35	2°0 9°0
Product % Olefin %	14.0-16.0	14.0-15.0	14.0-15.0	14.0-16.0 0.8 14.0-16.0 0.2 14.0-16.0 0.9	15.0-15.4 15.2-15.4 15.2-15.6 16.6-16.8	15.6 15.4 15.4 14.5 13.0	13.0 14.0
nd % C4H6	0-16.0 0.5-0.9	0.9-1.1	0.4-0.9	0.8-1.0	0 0 0 0 0 0 0	0.7 0.5-0.7 0.5-0.7 0.5	0.75 0.9
Feed % 01efin	14.0-16.0	14.0-16.0	14.0-15.0	14.0-15.0	15.1-15.8 15.2 15.2-15.8 17.0-17.2	15.0 15.6 15.6 15.0 14.0	13,4 14,9
-0 г. Кавотог 2	. 356		. 347	347 356	471-473 482-489 586-550 572-594	559 576-579 575-586 572-594 572-626	482 672–579
Temp. Fr. Reactor 1 Rea	į	356	347	320 356 356	471~482 478~486 536~545 572~581	559-561 576-590 572 672 672	482-518 572-579 572-590
Space Velocity Vol/Vol/hour	1000	2000-4000	3000-4000	460 1330 2670	0092	2600 3000 1000 600 260	2600 1600 1500
Catalyst	4788	3076	4788) 3076)	3076	4788) 3076)	4788)	3076